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<b>(21) International Application Number:</b> PCT/US98/01835 <b>(22) International Filing Date:</b> 30 January 1998 (30.01.98)  <b>(30) Priority Data:</b> 60/037,186      4 February 1997 (04.02.97)      US 09/010,776      22 January 1998 (22.01.98)      US  <b>(71) Applicant:</b> EASTMAN CHEMICAL COMPANY [US/US]; 100 North Eastman Road, Kingsport, TN 37660 (US).  <b>(72) Inventors:</b> TUSTIN, Gerald, Charles; 1065 Rotherwood Drive, Kingsport, TN 37660 (US). ZOELLER, Joseph, Robert; 2421 Rivermont Drive, Kingsport, TN 37660 (US). BROWNING, Horace, Lawrence, Jr.; 1113 Radcliffe Avenue, Kingsport, TN 37664 (US). SINGLETON, Andy, Hugh; 4137 Skyland Drive, Kingsport, TN 37664 (US).  <b>(74) Agent:</b> BLAKE, Michael, J.; P.O. Box 511, Kingsport, TN 37662-5075 (US).		<b>(81) Designated States:</b> CA, CN, JP, SG, European patent (AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE).  <b>Published</b> <i>With international search report.</i>
<b>(54) Title:</b> HETEROGENEOUS VAPOR PHASE CARBONYLATION PROCESS  <b>(57) Abstract</b>  Disclosed is a process wherein a mixture of methanol or a methanol source, a halide and carbon monoxide are contacted in the vapor phase with a supported catalyst comprising iridium and at least one second metal selected from ruthenium, molybdenum, tungsten, palladium, platinum and rhenium deposited on a catalyst support material.		

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- 1 -

HETEROGENEOUS VAPOR PHASE CARBONYLATION PROCESSReference to Related Application

This application claims the benefit of U.S.  
5 Provisional Application No. 60/037,186, filed  
February 04, 1997.

Field of the Invention

This invention pertains to a novel process for the  
10 production of acetic acid, methyl acetate or a mixture  
thereof by the carbonylation of methanol or methanol  
source. More specifically, this invention pertains to a  
process wherein a mixture of methanol or a methanol  
source, a halide and carbon monoxide are contacted in  
15 the vapor phase with a supported catalyst comprising  
iridium and at least one second metal selected from  
ruthenium, molybdenum, tungsten, palladium, platinum and  
rhenium deposited on a catalyst support material. The  
novel process provides a number of advantages, including  
20 improved production rates, with respect to known  
processes for the manufacture of acetic acid, methyl  
acetate or a mixture thereof from methanol.

Background of the Invention

25 Processes for the manufacture of acetic acid from  
methanol by carbonylation are operated extensively  
throughout the world. A thorough review of these  
commercial processes and other approaches to  
accomplishing the formation of acetyls from single  
30 carbon sources is described by Howard et al. in  
*Catalysis Today*, 18 (1993) 325-254. All commercial  
processes for the preparation of acetic acid by the  
carbonylation of methanol presently are performed in the  
liquid phase using homogeneous catalyst systems  
35 comprising a Group VIII metal and iodine or an iodine-

- 2 -

containing compound such as hydrogen iodide and/or methyl iodide. Rhodium is the most common Group VIII metal catalyst and methyl iodide is the most common promoter. These reactions are conducted in the presence of water to prevent precipitation of the catalyst. U.S. Patent 5,144,068 describes the inclusion of lithium in the catalyst system which allows the use of less water in the Rh-I homogeneous process. Iridium also is an active catalyst for methanol carbonylation reactions but normally provides reaction rates lower than those offered by rhodium catalysts when used under otherwise similar conditions. U.S. Patent 5,510,524 teaches that the addition of rhenium improves the rate and stability of both the Ir-I and Rh-I homogeneous catalyst systems. European Patent Application EP 0 752 406 A1 teaches that ruthenium, osmium, rhenium, zinc, cadmium, mercury, gallium, indium, or tungsten improve the rate and stability of the liquid phase Ir-I catalyst system. Generally, the homogeneous carbonylation processes presently being used to prepare acetic acid provide relatively high production rates and selectivity. However, heterogeneous catalysts offer the potential advantages of easier product separation, lower cost materials of construction, facile recycle and even higher rates.

Schultz, in U.S. Patent 3,689,533, describes the use of supported rhodium as a heterogeneous catalyst for the carbonylation of alcohols to carboxylic acids in the vapor phase in the presence of a halide promoter. In U.S. Patent 3,717,670, Schultz describes a similar supported rhodium catalyst in combination with promoters selected from Groups IB, IIIB, IVB, VB, VIB, VIII, (current notations: 11, 3, 4, 5, 6 and 8-10 respectively) lanthanide and actinide elements of the Periodic Table. Uhm, in U.S. Patent 5,488,143,

- 3 -

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support. In addition, the alcohol carbonylation process of Patent 3,772,380, described herein, is promoted, heterogeneous alcohol carbonylation is supported as a catalyst by EP 0 461 802 A2 describe the use of special carbonylation catalysts. European Patent Application EP 0 759 419 A1, published 26 February 26, 1997, after the filing date of Provisional Application No. 60/037,186 cited above, pertains to a process for the carbonylation of an alcohol and/or a reactive derivative thereof. EP 0 759 419 A1 discloses a carbonylation reactor wherein an alcohol is carbonylated in the liquid phase in the presence of a homogeneous catalyst system and the off gas from this first reactor is then mixed with addition- al alcohol and fed to a second reactor comprising a supported catalyst. The homogeneous catalyst system utilized in the first reactor comprises a halogen component and a Group VIII metal selected from rhodium and iridium. When the Group VIII metal is iridium, the homogeneous catalyst system also may contain an optional co-promoter selected from the group consisting of ruthenium, osmium, rhenium, cadmium, mercury, zinc, indium and gallium. The supported catalyst employed in the second reactor comprises a Group VIII metal selected

- 4 -

from the group consisting of iridium, rhodium and nickel, and an optional metal promoter on a carbon support. The optional metal promoter may be iron, nickel, lithium and cobalt. In the process exemplified  
5 in EP 0 759 419 A1, the conditions within the second carbonylation reactor zone are such that mixed vapor and liquid phases are present in the second reactor. The presence of a liquid phase component in the second reactor inevitably leads to leaching of the active  
10 metals from the supported catalyst which, in turn, results in a substantial decrease in the activity of the catalyst.

The literature contains several reports of the use of rhodium-containing zeolites as vapor phase alcohol  
15 carbonylation catalysts at one bar (100 kPa) pressure in the presence of halide promoters. The lead references on this type of catalyst are presented by Maneck et al. in *Catalysis Today*, 3 (1988), 421-429. Gelin et al., in *Pure & Appl. Chem.*, Vol 60, No. 8, (1988) 1315-1320,  
20 provide examples of the use of rhodium or iridium contained in zeolite as catalysts for the vapor phase carbonylation of methanol in the presence of halide promoter. Krzywicki et al., in *Journal of Molecular Catalysis*, 6 (1979) 431-440, describe the use of silica,  
25 alumina, silica-alumina and titanium dioxide as supports for rhodium in the halide-promoted vapor phase carbonylation of methanol, but these supports are generally not as efficient as carbon. Luft et al., in U.S. Patent 4,776,987 and in related disclosures,  
30 describe the use of chelating ligands chemically attached to various supports as a means to attach Group VIII metals to a heterogeneous catalyst for the halide-promoted vapor phase carbonylation of ethers or esters to carboxylic anhydrides. Evans et al., in U.S. Patent  
35 5,185,462, describe heterogeneous catalysts for halide-

- 5 -

promoted vapor phase methanol carbonylation based on noble metals attached to nitrogen or phosphorus ligands attached to an oxide support. Panster et al., in U.S. Patent 4,845,163, describe the use of rhodium-containing organopolysiloxane-ammonium compounds as heterogeneous catalysts for the halide-promoted liquid phase carbonylation of alcohols. Drago et al., in U.S. Patent 4,417,077, describe the use of anion exchange resins bonded to anionic forms of a single transition metal as catalysts for a number of carbonylation reactions including the halide-promoted carbonylation of methanol. Although supported ligands and anion exchange resins may be of some use for immobilizing metals in liquid phase carbonylation reactions, in general, the use of supported ligands and anion exchange resins offer no advantage in the vapor phase carbonylation of alcohols compared to the use of the carbon as a support for the active metal component.

Nickel on activated carbon has been studied as a heterogeneous catalyst for the halide-promoted vapor phase carbonylation of methanol, and increased rates are observed when hydrogen is added to the feed mixture. Relevant references to the nickel-on-carbon catalyst systems are provided by Fujimoto et al. In *Chemistry Letters* (1987) 895-898 and in *Journal of Catalysis*, 133 (1992) 370-382 and in the references contained therein. Liu et al., in *Ind. Eng. Chem. Res.*, 33 (1994) 488-492, report that tin enhances the activity of the nickel-on-carbon catalyst. Mueller et al., in U.S. Patent 4,918,218, disclose the addition of palladium and optionally copper to supported nickel catalysts for the halide-promoted carbonylation of methanol. In general the rates of reaction provided by nickel-based catalysts are lower than those provided by the analogous rhodium-based catalysts when operated under similar conditions.

- 6 -

Other single metals supported on carbon have been reported by Fujimoto et al. in *Catalysis Letters*, 2 (1989) 145-148 to have limited activity in the halide-promoted vapor phase carbonylation of methanol. The most active of these metals is Sn. Following Sn in order of decreasing activity are Pb, Mn, Mo, Cu, Cd, Cr, Re, V, Se, W, Ge and Ga. None of these other single metal catalysts are nearly as active as those based on Rh, Ir, Ni or the catalyst of the present invention.

A number of solid materials have been reported to catalyze the carbonylation of methanol without the addition of the halide promoter. Gates et al., in *Journal of Molecular Catalysis*, 3 (1977/78) 1-9, describe a catalyst containing rhodium attached to polymer bound polychlorinatedthiophenol for the liquid phase carbonylation of methanol. Current, in European Patent Application EP 0 130 058 A1, describes the use of sulfided nickel containing optional molybdenum as a heterogeneous catalyst for the conversion of ethers, hydrogen and carbon monoxide into homologous esters and alcohols. Smith et al., in European Patent Application EP 0 596 632 A1, describe the use of mordenite zeolite containing Cu, Ni, Ir, Rh, or Co as catalysts for the halide-free carbonylation of alcohols. Feitler, in U.S. Patent 4,612,387, describes the use of certain zeolites containing no transition metals as catalysts for the halide-free carbonylation of alcohols and other compounds in the vapor phase. Wegman, in U.S. Patent 5,218,140, describes the use of polyoxometallate anions containing Fe, Ru, Os, Co, Rh, Ir, Ni, Pd or Pt as catalysts for the halide-free carbonylation of alcohols and other compounds in the vapor phase. Although alcohol carbonylation without the use of halide promoters is very desirable, none of these halide-free



- 7 -

catalyst systems have rates comparable to those of catalysts promoted by halide.

Summary of the Invention

5 This invention pertains to a vapor phase process for the preparation of acetic acid, methyl acetate or carbon monoxide in the presence of a catalyst comprising iridium and at least one second metal selected from ruthenium, molybdenum, tungsten, palladium, platinum and rhodium deposited on a catalyst support material.

10 operation in the vapor phase distinguishes the present process from most of the prior art pertaining to methanol carbonylation wherein the reaction is carried out in the liquid phase utilizing a dissolved (homogeneous) catalyst.

15 The present invention provides a process for the preparation of acetic acid, methyl acetate or a mixture thereof which comprises the steps of:

(1) feeding a gaseous mixture comprising methanol, carbon monoxide, and a halide selected from chlorine, bromine, iodine and compounds thereof to a carbonylation zone which (i) contains a supported catalyst comprising iridium and at least one second metal selected from ruthenium, molybdenum, tungsten, palladium, platinum and rhodium deposited on a catalyst support material and (ii) is maintained

20 under carbonylation conditions of temperature and pressure; and

25 (2) recovering a gaseous product comprising acetic acid, methyl acetate or a mixture thereof from the carbonylation zone.

30 Our novel heterogeneous process operates entirely in the gas phase, i.e., none of the compounds or materials present in the carbonylation zone or reactor exists in a

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- 8 -

mobile liquid phase. Several unique advantages are bestowed by operating in this mode. The following are some of the advantages provided by operating with a heterogeneous catalyst in the vapor phase as compared to the operation of existing commercial, liquid phase, homogenous processes:

(1) Elimination of the difficult and costly separation of catalyst materials from the liquid product obtained from homogeneous processes.

(2) Removal of the rate limitations imposed by mass transfer of carbon monoxide into the reaction medium which pose an ultimate ceiling on the achievable rates in the liquid phase, homogeneous process. Avoiding such rate limitation allows greatly enhanced reaction rates to be achieved in gas phase operation.

(3) Operation at a much lower pressure, permitting the use of lower cost materials in plant construction.

In addition, iodine-containing materials are less corrosive in the vapor phase permitting the possible use of lower cost metallurgies.

Compared to the majority of prior art, heterogeneous processes which operate under conditions in which the reactants and/or products are present as liquids, the process of the present invention, by operating entirely in the vapor phase, eliminates catalyst dissolution (leaching from the catalyst support), which occurs in the known heterogeneous processes operating in the presence of liquid components.

Furthermore, the known operations in the presence of liquids are subject to the same problems associated with mass transfer of CO into a liquid reaction medium and the same upper limits associated with the homogeneous processes are observed. The process of the present invention provides faster rates and at lower pressures

- 9 -

as a consequence of operating with a very active catalyst in the vapor phase. Lastly, when compared to the known, rhodium-catalyzed, vapor phase processes, which represent the best operation known to date, the catalyst used in our novel process is considerably less expensive, yet performs at a comparable rate.

#### Description of the Invention

The catalyst for the process of the invention comprises a mixture of iridium and at least one second metal selected from ruthenium, molybdenum, tungsten, palladium, platinum and rhenium deposited on a catalyst support material. Examples of potentially useful support materials include carbon and the oxides or mixed oxides of silicon, aluminum, zinc, zirconium, or titanium. The preferred support materials are selected from the wide variety of carbon, activated carbon, and silicon oxide sources available commercially. Oxides of aluminum, or materials containing oxides of aluminum are least preferred.

Typically, the catalyst is prepared by (1) dissolving soluble compounds of iridium and the second metal (or metals) in an appropriate solvent; (2) combining the solution of step (1) with the support material; (3) slowly evaporating the solvent; and, optionally, (4) heating the dried catalyst from step (3) at elevated temperature in a stream of inert gas or in a vacuum. Variations in this method of catalyst preparation are well known to those skilled in the art and may be used.

The compound or form of iridium used to prepare the catalyst generally is not critical, and the catalyst may be prepared from any of a wide variety of iridium containing compounds. Indeed, iridium compounds containing myriad combinations of halide, trivalent

- 10 -

nitrogen, organic compounds of trivalent phosphorous, carbon monoxide, hydrogen, and 2,4-pentanedione, either alone or in combination, are available commercially and may be used in the preparation of the catalysts utilized in the present invention. In addition, the oxides of iridium may be used if dissolved in the appropriate medium. Based on their availability, cost, and high solubility in water (the preferred solvent medium) the preferred sources of iridium is one of its chlorides, such as the hydrated trichloride and any of the various salts of hexachloroiridate(IV). Use of either iridium trichloride or the hexachloroiridate complexes should be comparable on the basis of cost, solubility, and performance.

Likewise, the compound or form of the second metal compound used to prepare the catalyst generally is not critical, and the catalyst may be prepared using any of a wide variety of compounds containing ruthenium, molybdenum, tungsten, palladium, platinum or rhenium. A wide variety of compounds of these elements containing various combinations of halides, acetates, nitrates, trivalent nitrogen, organic compounds of trivalent phosphorous, carbon monoxide, hydrogen, and 2,4-pentanedione, either alone or in combination, are available commercially and may be used in the preparation of the catalysts utilized in the present invention. In addition, the oxides of these materials may be used if dissolved in the appropriate medium. However, the compound used to provide the second metal preferably is a water soluble form of the metal(s). Based on their availability and cost, the preferred sources are the various acetates, nitrates, and halides of the second metals. The most preferred source among these salts would be dictated by its solubility (preferably water solubility) which can vary widely across this list of

- 11 -

useful second components. For example, for the most preferred second promoters (Ru and Pd) the halides are generally available and quite soluble.

5 The content of iridium and the second component present on the catalysts can vary over a wide range, for example from 0.01 to 10 weight percent for each metal. However, the preferred catalysts contain 0.1 to 2 weight percent of each component. The second metal component of the catalysts utilized in the present invention preferably is ruthenium or molybdenum. The supported carbonylation catalysts which are particularly preferred 10 comprise (i) iridium and ruthenium or (ii) iridium and molybdenum deposited on a catalyst support material selected from carbon, activated carbon, and silicon oxide wherein the iridium and ruthenium each constitute 0.1 to 2 weight percent of the weight of the catalyst.

15 Although the methanol utilized in the process normally is fed as methanol, it can be supplied in the form of a combination of materials which generate methanol reactant. Examples of such combination of materials include (i) methyl acetate and water and (ii) dimethyl ether and water. In the operation of the process, both methyl acetate and dimethyl ether are 20 formed within the reactor and, unless methyl acetate is the desired product, they are recycled with water to the reactor where they are later consumed to form acetic acid.

25 Although the presence of water in the gaseous feed mixture is not essential when using methanol, the presence of some water is desirable to suppress formation of methyl acetate and/or dimethyl ether. When using methanol to generate acetic acid, the molar ratio of water to methanol can be 0:1 to 10:1, but preferably 30 is in the range of 0.01:1 to 1:1. When using an alternative source of methanol such as methyl acetate or 35

- 12 -

dimethyl ether, the amount of water fed usually is increased to account for the mole of water required for hydrolysis of the methanol alternative. Therefore, when using either methyl acetate or dimethyl ether, the mole ratio of water to ester or ether is in the range of 1:1 to 10:1, but preferably in the range of 1:1 to 3:1. In the preparation of acetic acid, it is apparent that combinations of methanol, methyl ester, and/or dimethyl ether are equivalent, provided the appropriate amount of water is added to hydrolyze the ether or ester to provide the methanol reactant.

When the process is operated to produce methyl acetate, no water should be added and dimethyl ether becomes the preferred feedstock. Further, when methanol is used as the feedstock in the preparation of methyl acetate, it is necessary to remove water. However, the primary utility of the process of the present invention is in the manufacture of acetic acid.

The gaseous mixture fed to the carbonylation zone also contains a halide component selected from chlorine, bromine, iodine and compounds thereof. The preferred halide components are selected from bromine, bromine compounds and, especially, iodine and iodine compounds which are gaseous under the carbonylation conditions of temperature and pressure. The halide components may be fed in various forms although the preferred forms are as an alkyl halide, especially as methyl halide, the hydrogen halide, or as the molecular halide, i.e.,  $I_2$ ,  $Br_2$ , or  $Cl_2$ . The molar ratio of methanol (or methanol equivalents) to halide or halide compound may vary substantially but typically is in the range of 1:1 to 10,000:1, with the preferred range being between 5:1 and 1000:1.

The carbon monoxide may be fed to the carbonylation zone either as purified carbon monoxide or as a mixture

- 13 -

of hydrogen and carbon monoxide. Although a small quantity of hydrogen may be useful in maintaining optimal catalyst activity, e.g., CO:hydrogen volume ratios of 99.5:0.5 to 95:5, the ratio of carbon monoxide to hydrogen, the presence of hydrogen, is believed to be of only minor importance. Thus, CO:hydrogen ratios of 100:0 to 25:75 all seem to be useful. However, the preferred CO:hydrogen ratios normally are in the range of 99:1 to 2:1.

The process of this invention is operated in the vapor phase and, therefore, is practiced at temperatures above the dew point of the product mixture, i.e., the temperature at which condensation occurs. However, since the dew point is a complex function of dilution (particularly with respect to non-condensable gases such as unreacted carbon monoxide, hydrogen, or inert diluent gas), product composition, and pressure, the process may still be operated over a wide range of temperatures, provided the temperature exceeds the dew point of the product effluent. In practice, this generally dictates a temperature range of 100 to 350°C, with temperatures in the range of 150 to 275°C being particularly useful.

As with temperature, the useful pressure range is limited by the dew point of the product mixture.

However, provided that the reaction is operated at a temperature sufficient to prevent liquefaction of the product effluent, a wide range of pressures may be used, e.g., pressures in the range of 0.1 to 100 bars absolute (bara) (10 to 10,000 kPa). The process preferably is carried out at a pressure in the range of 1 to 50 bara (100 to 5,000 kPa), most preferably, 3 to 30 bara (300 to 3,000 kPa).

- 14 -

Examples

The process of the present invention and the catalysts utilized in the process are further illustrated by the following examples.

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PREPARATION OF CATALYSTSCATALYST EXAMPLE 1

Iridium trichloride hydrate (418.9 mg) and ruthenium trichloride hydrate (275.7 mg) were dissolved  
10 in deionized water (30 mL). This mixture was added to 12 X 40 mesh activated carbon granules (20.0 g) having a BET surface area in excess of 800 m<sup>2</sup>/g contained in an evaporating dish. The mixture was heated on the steam bath with occasional stirring until it became free  
15 flowing and then transferred to a quartz tube measuring 106 cm long by 25 mm outer diameter. The quartz tube containing the mixture was placed in a three-element electric tube furnace so that the mixture was located in the approximate center of the 61 cm long heated zone of  
20 the furnace. Nitrogen (100 standard cubic centimeters per minute) was continuously passed through the catalyst bed, and the tube was heated from ambient temperature to 300°C over a 2 hour period, held at 300°C for 2 hours and then allowed to cool back to ambient temperature.  
25 The catalyst prepared in this manner (Catalyst I) contained 1.00 weight percent Ir and 0.49 weight percent Ru and had a density of 0.57 g per mL.

The preceding procedure was repeated three times: first, to prepare a catalyst containing only iridium as  
30 the active metal (Comparative Catalyst C-I); second, to prepare a catalyst containing only ruthenium (Comparative Catalyst C-II); and third, by substituting 303.5 mg (1.166 mmol) of rhodium trichloride trihydrate for the iridium trichloride hydrate, to prepare a



- 15 -

catalyst containing rhodium and ruthenium (Comparative Catalyst C-III).

#### CATALYST EXAMPLE 2

5       The procedure described in Example 1 was repeated except that 206.7 mg (1.166 mmol) of palladium chloride was substituted for the ruthenium trichloride hydrate to obtain a catalyst (Catalyst II) which contained 1.00 weight percent Ir and 0.49 weight percent Pd and had a  
10       density of 0.57 g per mL.

      The procedure of Catalyst Example 2 was repeated twice: once, to prepare a catalyst containing only palladium as the active metal (Comparative Catalyst C-IV); and second, by substituting 303.5 mg (1.166 mmol)  
15       of rhodium trichloride trihydrate for the iridium trichloride hydrate, to prepare a catalyst containing rhodium and palladium (Comparative Catalyst C-V).

#### CATALYST EXAMPLE 3

20       The procedure described in Example 1 was repeated except that 312.7 mg (1.166 mmol) of ammonium perrhenate ( $\text{NH}_4\text{ReO}_4$ ) was substituted for the ruthenium trichloride hydrate to obtain a catalyst (Catalyst III) which contained 1.00 weight percent Ir and 0.49 weight percent  
25       Re and had a density of 0.57 g per mL.

#### CATALYST EXAMPLE 4

      The procedure described in Example 1 was repeated except that 20 g of silica gel (Davison Chemical  
30       Company, Baltimore, Md, Grade 57, mesh 8) was substituted for carbon during catalyst preparation to produce a catalyst comprising 1.00 weight percent Ir and 0.49 weight percent Ru on silica (Catalyst IV).

- 16 -

CATALYST EXAMPLE 5

The procedure described in Example 1 was repeated except that 331.06 mg (1.166 mmol of tungsten) of ammonium tungsten oxide  $[(\text{NH}_4)_6\text{W}_{12}\text{O}_{41}] \cdot 5 \text{H}_2\text{O}$  was substituted for the ruthenium trichloride hydrate to obtain a catalyst (Catalyst V) which contained 1.08 weight percent Ir and 1.03 weight percent tungsten on activated carbon.

CATALYST EXAMPLE 6

The procedure described in Example 1 was repeated except that 205.85 mg (1.166 mmol of molybdenum) of ammonium molybdenum oxide  $[(\text{NH}_4)_6\text{Mo}_7\text{O}_{24}] \cdot 5 \text{H}_2\text{O}$  was substituted for the ruthenium trichloride hydrate to obtain a catalyst (Catalyst VI) which contained 1.05 weight percent Ir and 0.52 weight percent molybdenum on activated carbon.

COMPARATIVE CATALYST EXAMPLE C-6

The procedure described in Example 1 was repeated except that 1.153 g (4.665 mmol) of chromium acetate monohydrate was substituted for the ruthenium trichloride hydrate to obtain a catalyst (Comparative Catalyst C-VI) which contained 1.09 weight percent Ir and 0.29 weight percent chromium on activated carbon.

COMPARATIVE CATALYST EXAMPLE C-7

The procedure described in Example 1 was repeated except that 290.4 mg (1.166 mmol) of cobalt acetate tetrahydrate was substituted for the ruthenium trichloride hydrate to obtain a catalyst (Comparative Catalyst C-VII) which contained 1.08 weight percent Ir and 0.33 weight percent cobalt on activated carbon.

- 17 -

COMPARATIVE CATALYST EXAMPLE C-8

The procedure described in Example 1 was repeated except that 106.46 mg (1.166 mmol) of ferrous chloride was substituted for the ruthenium trichloride hydrate to obtain a catalyst (Comparative Catalyst C-VIII) which contained 1.09 weight percent Ir and 0.32 weight percent iron on activated carbon.

CARBONYLATION OF METHANOL

The reactor system consisted of a 800 to 950 mm (31.5 and 37 inch) section of 6.35 mm ( $\frac{1}{4}$  inch) diameter tubing constructed of Hastelloy alloy. The upper portion of the tube constituted the preheat and reaction (carbonylation) zones which were assembled by inserting a quartz wool pad 410 mm from the top of the reactor to act as support for the catalyst, followed sequentially by (1) a 0.7 g bed of fine quartz chips (840 microns), (2) 0.5 g of one of the catalysts prepared as described in the preceding examples, and (3) an additional 6 g of fine quartz chips. The top of the tube was attached to an inlet manifold for introducing liquid and gaseous feeds.

The six g of fine quartz chips acted as a heat exchange surface to vaporize the liquid feeds. Care was taken not to allow any liquid feeds to contact the catalyst bed at any time, including assembly, start-up, operation, and shutdown. The remaining lower length of tubing (product recovery section) consisted of a vortex cooler which varied in length depending on the original length of tubing employed and was maintained at approximately 0-5°C during operation.

The gases were fed using Brooks flow controllers and liquids were fed using a high performance liquid chromatography pump. The gaseous products leaving the reaction zone were condensed using a vortex cooler

- 18 -

operating at 0-5°C. The product reservoir was a tank placed downstream from the reactor system. The pressure was maintained using a modified Research control valve on the outlet side of the reactor system and the temperature of the reaction section was maintained using heating tape on the outside of the reaction system.

Feeding of hydrogen and carbon monoxide to the reactor was commenced while maintaining the reactor at a temperature of 240°C and a pressure of 17.2 bara (250 psia; 1,720 kPa). The flow rate of hydrogen was set at 25 standard cubic cm. per minute (cc/min) and the carbon monoxide flow rate was set at 100 cc/min. The reactor section was maintained under these conditions for 1 hour or until the temperature and pressure had stabilized (whichever was longer.) The high pressure liquid chromatography pump was then started, feeding a mixture consisting of 70 weight percent methanol and 30 weight percent methyl iodide at a rate of 10-12 g per hour. Samples of the liquid product were collected and analyzed periodically using gas chromatographic techniques.

#### CARBONYLATION EXAMPLE 1

The composition and weight of the samples taken periodically during the procedure described above in which Catalyst I was used are set forth in Table I wherein "Time" is the total time of operation (in hours) of the carbonylation commencing with the feeding of the methanol until a particular sample was taken. The values set forth below "MeI" (methyl iodide), "MeOAc" (methyl acetate), "MeOH" (methanol) and "HOAc" (acetic acid) are the weight percentages of each of those compounds present in the sample. The weight of each sample is given in grams.

- 19 -

TABLE I

	<u>Sample Number</u>	<u>Time</u>	<u>MeI</u>	<u>MeOAc</u>	<u>MeOH</u>	<u>HOAc</u>	<u>Sample Weight</u>
	1	3	22.11	26.97	28.03	5.57	33.5
5	2	5.5	22.34	27.58	27.74	6.45	33.3
	3	7.5	21.86	28.89	27.31	6.93	20.1
	4	10	23.0	31.22	24.65	8.0	33.5
	5	13	22.74	31.08	22.56	9.18	52.7
	6	15.5	21.71	28.83	26.18	7.27	31.1
10	7	19	22.48	31.66	24.29	8.61	42.7
	8	22	21.6	35.2	12.44	14.57	34.9
	9	25.5	22.36	36.13	13.59	13.51	45.5
	10	27	21.12	39.93	14.85	12.89	18.6

15 The rate of acetyl production based on the preceding  
 experiment utilizing Catalyst I is set forth in Table II  
 wherein Sample Number and Time values correspond to  
 those of Table I. "Acetyl Produced" is the amount  
 (millimoles) of methyl acetate and acetic acid produced  
 20 during each increment of Time calculated from the  
 formula:

$$\text{Sample Weight} \times 10 \times \frac{\text{Weight \% MeOAc}}{74} + \frac{\text{Weight \% AcOH}}{60}$$

25 "Production Rate" is the moles of Acetyl Produced per  
 liter of catalyst volume per hour during each increment  
 of Time (Time Increment), i.e., the time of operation  
 between samples. The formula for determining moles of  
 Acetyl Produced per liter of catalyst volume per hour  
 30 is:

$$\frac{\text{Acetyl Produced}}{0.5 \times \text{Time Increment}} \times 0.57$$

wherein 0.5 is the grams of catalyst used and 0.57 is  
 35 the density of the catalyst in g/mL.

- 20 -

TABLE II

	<u>Sample Number</u>	<u>Time</u>	<u>Acetyl Produced</u>	<u>Production Rate</u>
	1	3	153.2	58.2
5	2	5.5	159.9	72.9
	3	7.5	101.7	58.0
	4	10	186.0	84.8
	5	13	302.0	114.7
	6	15.5	158.8	72.4
10	7	19	244.0	79.5
	8	22	250.8	95.3
	9	25.5	324.6	105.7
	10	27	140.4	106.7

Repetition of the above catalyst preparation and carbonylation procedures demonstrated a period of 3-8 hours during which the reaction rate continued to increase significantly (an induction period). After completion of the induction period, acetyl production rates averaged 90 moles/L-hour for 4 separate carbonylation experiments. The average acetyl production rate for the carbonylation experiment described herein was 94 moles/L-hour after the induction period.

CARBONYLATION EXAMPLES 2-6 AND COMPARATIVE CARBONYLATION  
EXAMPLES C-1 - C-9

Catalysts II-VI and Comparative Catalysts C-I - C-IX were utilized in the carbonylation of methanol according to the above-described procedure. The Production Rate, i.e., the moles of Acetyl Produced per liter of catalyst volume per hour, provided by each of Catalysts II-VI and Comparative Catalysts C-I - C-IX is shown in Table III.

- 21 -

TABLE III

	<u>Carbonylation</u> <u>Example</u>	<u>Catalyst</u>	<u>Production</u> <u>Rate</u>
5	C-1	C-I	52
	C-2	C-II	20
	C-3	C-III	34
	2	II	86
	C-4	C-IV	2
10	C-5	C-V	61
	3	III	82
	4	IV	81
	5	V	93
	6	VI	123
15	C-6	C-VI	4
	C-7	C-VII	40
	C-8	C-VIII	3

- 22 -

CLAIMS

We claim:

1. Process for the preparation of acetic acid, methyl acetate or a mixture thereof which comprises the steps of:

(1) feeding a gaseous mixture comprising methanol, carbon monoxide, and a halide selected from chlorine, bromine, iodine and compounds thereof to a carbonylation zone which (i) contains a supported catalyst comprising iridium and at least one second metal selected from ruthenium, molybdenum, tungsten, palladium, platinum and rhenium deposited on a catalyst support material and (ii) is maintained under carbonylation conditions of temperature and pressure; and

(2) recovering a gaseous product comprising acetic acid, methyl acetate or a mixture thereof from the carbonylation zone;

wherein none of the compounds present in the carbonylation zone exists in a mobile liquid phase.

2. Process according to Claim 1 wherein the halide is selected from iodine, hydrogen iodide and methyl iodide and the carbonylation zone is maintained at a temperature of 100 to 350°C and a pressure of 100 to 5,000 kPa.

3. Process according to Claim 2 wherein the gaseous mixture includes water in an amount which gives a water:methanol mole ratio of 0.01:1 to 1:1.

4. Process according to Claim 1 wherein the catalyst comprises 0.1 to 2 weight percent each of iridium and the second metal and the catalyst support material is



- 23 -

selected from carbon, activated carbon, and silicon oxide.

- 5      5.    Process according to Claim 1 wherein the supported catalyst comprises iridium, ruthenium and a catalyst support material.
- 10     6.    Process according to Claim 5 wherein the halide is selected from iodine, hydrogen iodide and methyl iodide and the carbonylation zone is maintained at a temperature of 100 to 350°C and a pressure of 100 to 5,000 kPa.
- 15     7.    Process according to Claim 6 wherein the gaseous mixture includes water in an amount which gives a water:methanol mole ratio of 0.01:1 to 1:1.
- 20     8.    Process according to Claim 5 wherein the catalyst comprises 0.1 to 2 weight percent each of iridium and ruthenium and the catalyst support material is selected from carbon, activated carbon, and silicon oxide.
- 25     9.    Process for the preparation of acetic acid which comprises the steps of:
- 30        (1)    feeding a gaseous mixture comprising methanol, water, carbon monoxide, and a halide selected from iodine, hydrogen iodide and methyl iodide to a carbonylation zone which (i) contains a supported catalyst comprising iridium, ruthenium and a catalyst support material selected from carbon, activated carbon, and silicon oxide and (ii) is maintained at a temperature of 150 to 275°C and a pressure of 300 to 5,000 kPa; and
- 35        (2)    recovering a gaseous product comprising acetic acid from the carbonylation zone;

- 24 -

wherein none of the compounds present in the carbonylation zone exists in a mobile liquid phase.

- 5 10. Process according to Claim 9 wherein the water:methanol mole ratio in the gaseous feed to the carbonylation zone is 0.01:1 to 1:1 and the catalyst comprises 0.1 to 2 weight percent each of iridium and ruthenium.
- 10 11. Process according to Claim 1 wherein the supported catalyst comprises iridium, molybdenum and a catalyst support material.
- 15 12. Process according to Claim 11 wherein the halide is selected from iodine, hydrogen iodide and methyl iodide and the carbonylation zone is maintained at a temperature of 100 to 350°C and a pressure of 100 to 5,000 kPa.
- 20 13. Process according to Claim 12 wherein the gaseous mixture includes water in an amount which gives a water:methanol mole ratio of 0.01:1 to 1:1.
- 25 14. Process according to Claim 11 wherein the catalyst comprises 0.1 to 2 weight percent each of iridium and molybdenum and the catalyst support material is selected from carbon, activated carbon, and silicon oxide.
- 30 15. Process for the preparation of acetic acid which comprises the steps of:
- 35 (1) feeding a gaseous mixture comprising methanol, water, carbon monoxide, and a halide selected from iodine, hydrogen iodide and methyl iodide to a carbonylation zone which (i) contains a supported catalyst comprising iridium, molybdenum and a

- 25 -

catalyst support material selected from carbon, activated carbon, and silicon oxide and (ii) is maintained at a temperature of 150 to 275°C and a pressure of 300 to 5,000 kPa; and

- 5 (2) recovering a gaseous product comprising acetic acid from the carbonylation zone;

wherein none of the compounds present in the carbonylation zone exists in a mobile liquid phase.

- 10 16. Process according to Claim 15 wherein the water:methanol mole ratio in the gaseous feed to the carbonylation zone is 0.01:1 to 1:1 and the catalyst comprises 0.1 to 2 weight percent each of iridium and molybdenum.

15

## INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 98/01835

A. CLASSIFICATION OF SUBJECT MATTER IPC 6 C07C51/12 C07C53/08 C07C67/40		
According to International Patent Classification (IPC) or to both national classification and IPC		
B. FIELDS SEARCHED		
Minimum documentation searched (classification system followed by classification symbols) IPC 6 C07C		
Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched		
Electronic data base consulted during the international search (name of data base and, where practical, search terms used)		
C. DOCUMENTS CONSIDERED TO BE RELEVANT		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	DD 296 275 A (AKADEMIE DER WISSENSCHAFTEN) 28 November 1991 see page 2, line 24 - page 2, line 30; claims 1-5	1-14
A	--- US 5 510 524 A (GARLAND CARL S ET AL) 23 April 1996 cited in the application see claim 1	1-16
A	--- US 3 772 380 A (PAULIK F ET AL) 13 November 1973 cited in the application see abstract see example 5 -----	1-16
<input type="checkbox"/> Further documents are listed in the continuation of box C. <input checked="" type="checkbox"/> Patent family members are listed in annex.		
* Special categories of cited documents : * A* document defining the general state of the art which is not considered to be of particular relevance * E* earlier document but published on or after the international filing date * L* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) * O* document referring to an oral disclosure, use, exhibition or other means * P* document published prior to the international filing date but later than the priority date claimed * T* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention * X* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone * Y* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art. * &* document member of the same patent family		
Date of the actual completion of the international search  26 March 1998		Date of mailing of the international search report  09.04.98
Name and mailing address of the ISA European Patent Office, P.B. 5818 Patentlaan 2 NL - 2280 HV Rijswijk Tel. (+31-70) 340-2040, Tx. 31 651 epo nl, Fax: (+31-70) 340-3016		Authorized officer  Arias-Sanz, J

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Information on patent family members

International Application No

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